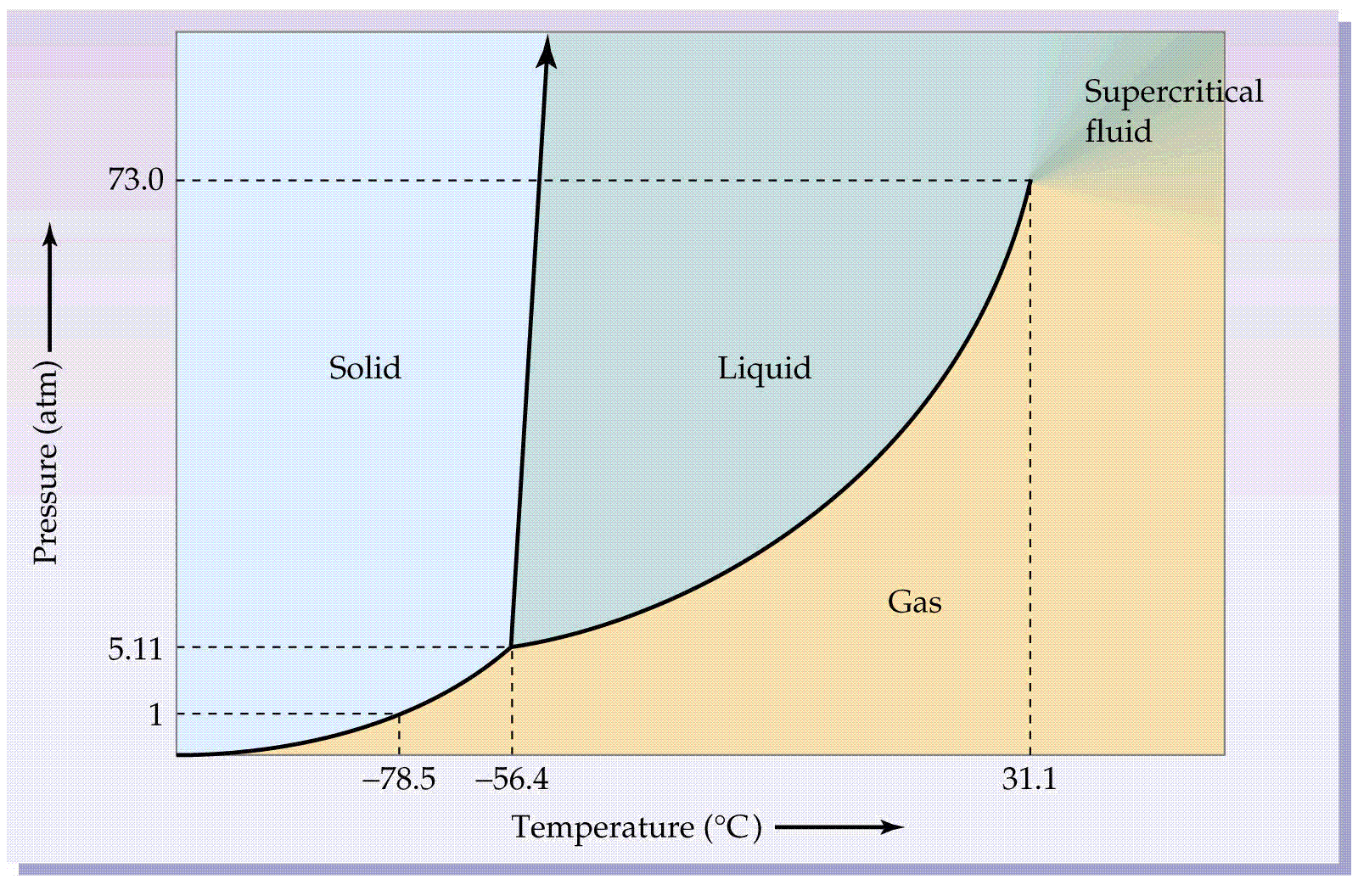
**Read This!**

A ***phase*** is any part of a system that which all physical properties of a material are uniform. A ***state of matter*** is one of the distinct forms that [matter](http://en.wikipedia.org/wiki/Matter) takes on. There are four observable states of matter in everyday life: *solid, liquid, gas* and *plasma*. ***Plasma***, which is ionized gas, is actually the most common visible state of matter in the universe as the stars are composed of plasma. On earth, every time you see a lightning bolt you are seeing plasma. Historically, the distinction is made based on qualitative differences in the properties of volume and shape. A ***solid*** has a fixed volume and a fixed shape; A ***liquid*** has a fixed volume and a variable shape; A ***gas*** has both a variable volume and a variable shape.

**Model 1 – Phase Diagram for Carbon Dioxide**



1. a. Model 1 is a Phase Diagram for what compound?

b. What is the formula for this compound?

2. Name the x-axis and y-axis in Model 1.

3. In what units are these two variables measured on Model 1?

4. a. Are the two axes equidistant (equally spaced) on Model 1?

b. Explain why or why not.

5. Name the three states of matter separated by solid lines on Model 1.

6. On Model 1, what do the three solid lines represent?

7. Trace over the solid line separating Solid / Gas in red. Trace over the solid line separating Liquid / Gas in blue and trace over the solid line separating Solid / Liquid in green.

8. The process (change of state) of ***condensation*** occurs when a gas changes into a liquid. Draw and label a blue arrow across the solid line showing the direction of *condensation*. The process of ***evaporation*** occurs when a liquid changes into a gas. Draw and label a blue arrow across the solid line showing the direction of *evaporation*.

9. The process of ***freezing*** occurs when a liquid changes into a solid and the reverse process is ***melting***. Draw and label green arrows showing these two phase changes.

10. The process of ***deposition*** occurs when a gas changes into solid and the reverse process is ***sublimation***. Draw and label red arrows showing these two phase changes.

11. Circle on Model 1 where the three solid lines meet. What is this junction called?

12. What is the purpose of the dashed lines on Model 1?

13. Use Model 1 to determine what state of matter CO2 isat -78.5C and 73.0 atm.

14. Use Model 1 to determine what state of matter CO2 isat 0C and 70.0 atm.

15. Use Model 1 to determine what state of matter CO2 isat 31.1C and 1 atm.

16. At a constant 10 atm, list all the states of matter and processes that CO2 undergoes when heated from

-78.5C to 35C.

17. At a constant 1atm, list all the states of matter and processes that CO2 undergoes when cooled from

30C to -100C.

18. At a constant -60C, list all the states of matter and processes that CO2 undergoes when depressurized

from 6 atm to 0.25 atm.

19. At -55.4C, explain all the states of matter and processes (changes of state) that CO2 undergoes when pressurized from 1 atm to 80 atm.

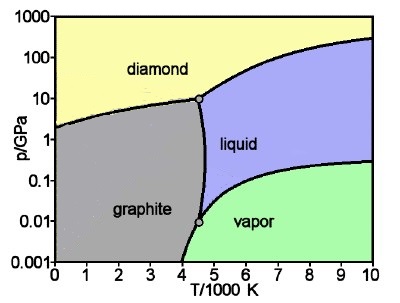
20. The ***normal boiling point*** of a substance occurs at the temperature in which that substance boils at 1 atm of pressure. The ***normal freezing point*** of a substance occurs at the temperature in at which that substance freezes at 1 atm. Examine Model 1 where 1 atm of pressure is on the solid line separating Solid / Gas. Suggest a name for this point.

**Read This!**

Allotropes are two or more forms of the same element in the same physical state (solid, liquid, or gas) that differ from each other in their physical, and sometimes chemical, properties. The most notable examples of allotropes are found in groups 14, 15, and 16 of the [periodic table](http://science.jrank.org/pages/5106/Periodic-Table.html). Gaseous [oxygen](http://science.jrank.org/pages/4970/Oxygen.html), for example, exists in three allotropic forms: monatomic oxygen (O), a diatomic [molecule](http://science.jrank.org/pages/4414/Molecule.html) (O2), and in a triatomic molecule known as [ozone](http://science.jrank.org/pages/4972/Ozone.html) (O3).

A striking example of differing physical properties among allotropes is the case of [carbon](http://science.jrank.org/pages/1203/Carbon.html). Solid carbon exists in two allotropic forms: [diamond](http://science.jrank.org/pages/2050/Diamond.html) and graphite. Diamond is the hardest naturally occurring substance and has the highest melting point of any element. In contrast, graphite is a very soft material, the substance from which the "lead" in lead pencils is made.

**Model 2 – Phase Diagram for Carbon**



21. a. Model 2 is a Phase Diagram for what element?

b. What is the symbol for this element?

22. Name the x-axis and y-axis in Model 2.

23. a. In what units are these two variables measured on Model 2?

b. What do these symbols stand for?

24. a. Are the two axes equidistant (equally spaced) on Model 2?

b. Explain why or why not.

25. a. Compare Model 1 to Model 2. On Model 2, which section do you think corresponds to the “gas” section on Model 1?

b. Define this word.

26. On Model 2, the phase section of solid is divided into what sections?

27. a. Use Model 2 to determine what state of matter carbon is in at 3000K and 100GPa.

b. If solid, give specific allotrope.

28. a. Use Model 2 to determine what state of matter carbon is in at 3000K and 0.01GPa.

b. If solid, give specific allotrope.

29. a. Use Model 2 to determine what state of matter carbon is in at 8000K and 1GPa.

b. If solid, give specific allotrope.

30. a. Use Model 2 to determine what state of matter carbon is in at 9000K and 0.01GPa.

b. If solid, give specific allotrope.

31. At what pressure and temperature is the triple point for carbon?

32. At a constant 0.1GPa, list all the states of matter and processes that carbon undergoes when heated from 3000K to 10,000K.

33. At a constant 0.002GPa, list all the states of matter and processes that carbon undergoes when cooled from 5000K to 2000K.

34. At a constant 4100K, list all the states of matter and processes that carbon undergoes when depressurized from 0.1GPa to 0.001GPa?

35. At a constant 7500K, list all the states of matter and processes that carbon undergoes when pressurized from 0.01GPa to 100GPa?

Extra Credit

38. Using Model 2, describe carbon at the point of 4.5K and 10 GPa.